** TENNESSEE STATE UNIVERSITY**

**COLLEGE OF LIFE & PHYSICAL SCIENCE**

**DEPARTMENT OF CHEMISTR**

**CHEM 2011-80: Organic Chemistry I Laboratory**

**Fall 2019**

**Instructor’s Name:** Mohammad Karim

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**Office Hours: MWF**: 10:00-11:00 AM

**Textbook:** “Macroscale and Microscale Organic Experiments”, 5th edition, Kenneth L. Williamson; Brooks Cole, 2006 [ISBN-13: 978-0618590674]

**Class:** CHEM 2011-80, **R (**5:30 – 8:30), Boswell 224

**Prerequisite:** **General Chemistry II (CHEM 1120 and 1121) – Lecture and Lab**

**Objective:** This course is designed or chemistry majors, pre-professionals, biology and other scientific areas. Its aims are to expose students to some of the important experimental techniques, equipment and reactions in organic chemistry, and to illustrate the chemical reactions and properties of molecules discussed in lecture.

**Course Description**: This course exposes students to some of the important experimental techniques, equipment and reactions in organic chemistry. In addition, the course illustrates the chemical reactions and properties of molecules discussed in lecture. Students will have hands-on experiences with purification methods such as: crystallization, distillation, extraction, TLC, column chromatography etc. The course also involves organic synthesis, purification and identification using melting points and spectroscopic techniques.

**Course Competencies:** Upon completion of this course, the student should have obtained the following competencies:

* A thorough understanding of the common isolation and purification techniques used in organic chemistry
* A working knowledge of the important spectroscopic techniques used for characterization and identification of organic molecules
* A working knowledge of the equipment and techniques used in performing simple organic synthesis
* Analyze mixture of compounds by TLC
* Set up and perform simple and fractional distillation and learn the efficiency of both techniques
* A thorough understanding of the scientific writing process

**GRADING: 90 – 100 = A, 80 – 89 = B, 70 – 79 = C, 60 = 69 = D, <60 = F**

**Notebook grade including Lab Performance = 40% of total grade**

**Quizzes = 10% of total grade**

**Formal report (1) = 20% of total grade**

**Workstation Organization/Tidiness and Teamwork = 10% of total grade**

**Final Exam = 20% of total grade**

**Lab Performance includes:**

* Preparation
* Participation in actual experiment
* Results

**LABORATORY REPORTS AND DUE DATES**:

**LATE LABORATORY REPORT:** Any laboratory report not turned in on or before the due date will be assessed a 5-point penalty per day. Reports will not be accepted if they are past due more than two days.

**OUTLINES:** A detail outline of the experimental procedure for each experiment must be completed before coming to class. The outline will help you complete the lab faster and also help you better understand the experiment.

**ATTENDANCE AND EXPECTATIONS:** Students are expected to regularly attend all classes in this course for credit and to complete all required work in this course. Students are expected to read the *Undergraduate Catalog* for further instructions on attendance policy.

**Department Policy on Make-Up Labs Due to University Closings: (**10-experiment semester)

* Missed Day 1: Since there is an extra day in the syllabus for review, this day will be used for make-up of that experiment. There will be 10 experiments for evaluation purposes.
* Missed Day 2: This lab will be made-up on Fridays. The specific time for each lab class will be forwarded to the students by the lab professor via the department chair. There will be 10 experiments for evaluation purposes.
* Missed Day 3: This lab will be made-up on Fridays. The specific time for each lab class will be forwarded to the students by the lab professor via the department chair. There will be 10 experiments for evaluation purposes.
* Missed Day 4: There will be no make-up. There will be 9 experiments for evaluation.
* Additional Missed Days: There will be no make-up. The number of experiments for evaluation purposes will be reduced by one for each additional missed day.

**Policy on Academic Misconduct, Cheating and Plagiarism:** In accordance with the university’s policy on academic and classroom misconduct found in the catalog, cheating will not be tolerated in this course and a zero-tolerance policy regarding cheating will be followed throughout the course. A student who is caught cheating or attempting to cheat will be given a zero (F) for that particular assignment/test/quiz for the first offense. If a student is caught cheating a second time, that student will be given an overall grade of “F” for the course. To this end, the following classroom policies will be in effect and enforced.

* Cell phones and any other electronic devices (including smart watches) that connect to wireless networks will not be permitted during any exam or quiz. These devices may not be on your desk during an exam or quiz and must be stored in your bag or purse and/or turned off. Calculators may be used, only if the questions on the exam or quiz warrant their use.
* Once an exam or quiz period has started, you will not be permitted to leave to go to the restroom during the exam period. Please be sure to use the restroom before coming to class. Exceptions will only be made for those with documented medical needs.
* No outside materials may be used during an exam or quiz. Any necessary materials (*i.e.* periodic table, equations & constants, scratch paper, *etc.*) will be provided for you.
* Sunglasses and hats may not be worn during an exam or quiz period.
* The use of headphones and/or earbuds during an exam or quiz is strictly prohibited.
* Duplication or copying of homework assignments will result in a score of zero (F) for each student submitting a copied homework assignment.

**TSU Disability Accommodation Statement:** TSU is committed to creating inclusive learning environments and providing all students with opportunities to learn and excel in their course of study. Any student with a disability or condition which might interfere with his/her class performance or attendance may arrange for reasonable accommodations by visiting the Office of Disability Services (ODS). ODS is located in Kean Hall, room 131 and can be reached at 963-7400 or [www.tnstate.edu/disabilityservices](https://email.tnstate.edu/owa/redir.aspx?C=gk6WOH_1TE-MCLQNo_mn52fQIPFZzNMIw444dBa7_m0A7UvXztod9aW6iBa4gjigMROwMmBdzho.&URL=http%3a%2f%2fwww.tnstate.edu%2fdisabilityservices) .  You will be required to speak with ODS staff and provide documentation of the need for an accommodation.  If you qualify for an accommodation you will be provided with a document stating what type of classroom accommodations are to be made by the instructor.  It is your responsibility to give a copy of this document to the instructor **as soon as you receive it**.  Accommodations will only be provided **AFTER** the instructor receives the accommodation instructions from ODS; accommodations are not retroactive.  You must follow this process for each semester that you require accommodations.

**TSU Sexual Misconduct, Domestic/Dating Violence, Stalking Statement:** TSU recognizes the importance of providing an environment free of all forms of discrimination and sexual harassment, including sexual assault, domestic violence, dating violence, and stalking.  If you (or someone you know) has experienced or is experiencing any of these incidents, there are resources to assist you in the areas of accessing health and counseling services, providing academic and housing accommodations, and making referrals for assistance with legal protective orders and more. Please be aware that most TSU employees, including faculty and instructors, are “responsible employees”, meaning that they are required to report incidents of sexual violence, domestic/dating violence or stalking.   **This means that if you tell me about a situation involving sexual harassment, sexual assault, dating violence, domestic violence, or stalking, I must report the information to the Title IX Coordinator.**  Although I have to report thesituation, you will still have options about how your situation will be handled, includingwhether or not you wish to pursue a formal complaint.  Our goal is to make sure you areaware of the range of options available to you and have access to the resources youneed.    
You are encouraged to contact TSU’s Title IX Coordinator to report any incidents of sexual harassment, sexual violence, domestic/dating violence or stalking.  The Title IX coordinator is located in the Office of Equity and Inclusion, McWherter Administration Building, Ste. 260 and can be reached at 963-7494 or 963-7438.  For more information about Title IX and TSU’s SART or policies and procedures regarding sexual, domestic/dating violence and stalking please visit:  [www.tnstate.edu/equity](https://email.tnstate.edu/owa/redir.aspx?C=gk6WOH_1TE-MCLQNo_mn52fQIPFZzNMIw444dBa7_m0A7UvXztod9aW6iBa4gjigMROwMmBdzho.&URL=http%3a%2f%2fwww.tnstate.edu%2fequity). If you wish to speak to someone confidentially, who is not required to report, you can contact the TSU Counseling Center, located in the basement of Wilson Hall, at 963-5611 or TSU Student Health Services, located in the Floyd Payne Campus Center room 304, at 963-5084.  You may also contact the following off campus resources:  Sexual Assault Center of Nashville at 1-800-879-1999 or [www.sacenter.org](https://email.tnstate.edu/owa/redir.aspx?C=gk6WOH_1TE-MCLQNo_mn52fQIPFZzNMIw444dBa7_m0A7UvXztod9aW6iBa4gjigMROwMmBdzho.&URL=http%3a%2f%2fwww.sacenter.org) or the Tennessee Coalition to End Domestic & Sexual Violence at 615-386-9406 or [www.tncoalition.org](https://email.tnstate.edu/owa/redir.aspx?C=gk6WOH_1TE-MCLQNo_mn52fQIPFZzNMIw444dBa7_m0A7UvXztod9aW6iBa4gjigMROwMmBdzho.&URL=http%3a%2f%2fwww.tncoalition.org).

**TSU Harassment and Discrimination Statement:** Tennessee State University is firmly committed to compliance with all federal, state and local laws that prohibit harassment and discrimination based on race, color, national origin, gender, age, disability, religion, retaliation, veteran status and other protected categories.  TSU will not subject any student to discrimination or harassment and no student shall be excluded from participation in nor denied the benefits of any educational program based on their protected class.  If a student believes they have been discriminated against or harassed because of a protected class, they are encouraged to contact the Office of Equity and Inclusion at McWherter Administration Building, Ste. 260, 615-963-7494 or 963-7438, [www.tnstate.edu/equity](https://email.tnstate.edu/owa/redir.aspx?C=gk6WOH_1TE-MCLQNo_mn52fQIPFZzNMIw444dBa7_m0A7UvXztod9aW6iBa4gjigMROwMmBdzho.&URL=http%3a%2f%2fwww.tnstate.edu%2fequity).

**SAFETY GOGGLES: Safety Goggles are required in this course. No student will be allowed to work in the laboratory without a pair of Safety Goggles. Safety Goggles must be worn at all times while in the laboratory.**

**TSU Statement of Disability Policy for Students**

It is the policy of TSU to accommodate students with disabilities, pursuant to federal law, state law, and the University’s commitment to equal educational opportunities. Any student with a disability who needs accommodations, for example in seating placement or in arrangements for examinations, should inform the instructor at the beginning of the course. Students with disabilities are encouraged to contact the Office of Disabled Student Services, which is located in Kean Hall, Room 117, Telephone 963-7400 or 963-7872.

**EXPERIMENTS:**

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| |  |  |  | | --- | --- | --- | | **Week** | **Reading Assignments** | **Experiments** | | 1 | Check-In Procedures. Chapter 1: Introduction, Laboratory notebook writing and related calculation. Chapter 2: Laboratory Safety and waste disposal. Safety Quiz | | | 2 | Chapter 3: Melting Points and Boiling Points (pp 38-58) | Expt. 1: Calibration of Thermometer (p 51)  Expt 2: Microscale procedure for pure Urea and Cinnamic Acid (p-52)  Expt 3: Melting Points of Urea-Cinnamic Acid Mixtures (p 53)  Expt. 4: Melting Points of Unknowns (p 53) | | 3 | Chapter 4: Recrystallization (pp 61-87) | Expt.1. Solubility Tests (p 79)  Expt. 2: Microscale procedure for Phthalic Acid (p 81) | | 4 | Chapter 5: Distillation (pp 88-104) | Expt 2A: Fractional Distillation of Cyclohexane andToluene Mixture (p-98)  Expt. 6: Fractional Distillation of Unknowns (103) | | 5 | Chapter 5: Continue | Expt 1: Separation of a Carboxylic Acid, a phenol, and a Neutral Substance (p-151) | | 6 | Chapter 7: Extraction (pp 135-171) | Expt 6: Extraction of Caffeine from Tea (p-162-164) | | 7 | Hand out– Aldehyde Functional Groups | Tollens Test – “Silver mirror experiment” | | 8 | Chapter 8: Thin Layer Chromatography: Analysis of Analgesics and Isolation of Lycopene from Tomato Paste (pp 172-193) | Expt 1: Analgesics (p-183) | | 9 | Chapter 9: Column Chromatography | Expt 4: Chromatography of the Ferrocene and Acetylferrocene (p 210-211) | | 10 | Chapter 16: The SN2 Reaction- 1-Bromobutane (pp 325-332) | Expt 1: Synthesis of 1-Bromobutane (p 327) | | 11 | Chapter 19: Alkenes from Alcohol. Cyclohexene from Cyclohexanol (pp 351-356) | Expt: 1: The Preparation of Cyclohexene (p 352) | | 12 | Review session and Check-Out | | | 13 | Final Exam. Lab reports and Lab notebooks will be collected for grading. | | |

**GENERAL LAB DIRECTIONS**

**A. Planning and Efficiency**

Plan your lab work ahead of time. This is essential if you expect to complete an experiment within the allotted time. Furthermore, you will have some understanding of what you are doing.

1. Read the experiment and all reference material before coming to lab. (Prepare a complete outline of the experiment).
2. Make good use of your time. When you are waiting for something to dry, or waiting for something to cool, etc., start on or prepare for the next step or experiment. Learn to do more than one thing at a time.

3. Listen for announcements from your T.A. and/or instructor. Periodically check the lab blackboard for notices.

1. Do not be afraid to ask the instructor questions about the experiment. They (we) are here to help you. At the same time, you should be prepared for them to ask you questions about the experiment or technique.

5. Every student in this course is required to do his or her own work: this includes all experimental work, calculations, and reports.

1. Be sure to label all solutions or chemicals in your desk so you do not mix them up. Do not rely on your memory. Use pencil to label. Do not use tape or ink.
2. Upon entering the lab, you will be assigned to a specific locker and work area. It’s a good idea to keep a towel in your locker. It is your responsibility to keep your work area, adjacent sink, and floor clean. Failure to do so will affect your grade. You are responsible for, together with your fellow students, for the cleanliness of community equipment - balances, hoods, and reagent tables. Again failure to assist in keeping the laboratory clean will lower your grade.

**B. General Comments Regarding Notebook**

1. Your notebook is the only record you have of your work in the lab.

2. You must use a bound composition notebook. Loose pages, three-ring notebooks, or spiral notebooks are not accepted.

3. The pages in the notebook should be numbered. The first 6 pages should be an up-to-date “Table of Contents” so that any experiment can be located quickly.

4. All entries in the notebook must be in ink. A mistake is not erased, but crossed out with a single line and replaced with the correction written above.

5. All data obtained in the laboratory are recorded in the notebook at the time the data are obtained. **Recording data on loose scraps of paper is not acceptable.**

6. Your notebook must be sufficiently neat that someone else (especially the instructor) can read and understand it. Use plenty of space. Remember, you will have to extract data from your notebook to calculate results and more importantly write reports.

**C. Format of Notebook:**

**Must write your Name, Course Name, and Instructor’s Name on front page of the notebook**

1. Reserve pages 1-4 for Table of Contents. Keep Table of Contents current.

2. Start Experiment 1 at the top of page **5**.

3. For each new experiment, begin at the top of a new right-hand page.

4. For each experiment, begin with the following items:

**I**. **Title of the experiment**, date of experiment

**Equation of the Experiment**

**Table of reagents**

**II**. A **brief** (3-4 sentences) statement of introduction, purpose, and principles on which the experiment(s) is based. Relevant balanced chemical equations should appear here when applicable.

**III**. An outline of the procedure to be followed.

**IV During experiment:** A complete and neat lists of data taken during the experiment. Data must be properly identified and labeled and presented in a logical order. Use appropriate structural drawings when necessary. Use tables where appropriate.

**V**. All other observations that you make during the course of the experiment should also be recorded. Examples include: color changes, changes you make in the general procedure, observed errors such as adding to much solvent or solute or too much of a certain reagent. If you are in doubt as to if the observation is relevant, make a note of it in your notebook. Do not trust your memory!

**Sections I, II and III above must be completed before coming to the lab. This work will credit you 10 % of Notebook grade. This will save significant time during the lab period.**

The Chemistry Laboratory Formal Report

There is not one absolutely correct format or method for writing up a laboratory report. The format given below is for a published experimental study and is the format that is expected in general. The goal is to "include everything necessary in the shortest possible way." Standard English is required, and everything that is written must be scientifically sound and logical.

1. **Title Page:**

This should as specific as possible. State the actual substances used and the method: Determination of the Molar Mass of Acetone by the Dumas Method.

1. Name:
2. Date Experiment Started and Date Finished
3. Date Submitted
4. **Abstract**

Put on the next page. In general, summarize the results of the experiment and state any significant conclusions. Specifically:

1. State the method used, and actual experimental values that you obtained with uncertainties
2. If there are literature values, put these here as well; (this is not necessary when actually publishing, but it is good practice for in-house reports, unless there are many literature values).
3. Put statement comparing your results with the literature values. If there is disagreement exceeding experimental error, state a plausible reason why there is a difference.
4. **Introduction**
5. State the purpose of the experiment.
6. Give a brief outline of the theory behind the experiment as well as any pertinent definitions. This includes the "working equation(s)" that will be used in treating your experimental data and the actual experiment quantities to be measured.
7. Make sure that all symbols are defined in any equations that are used.
8. **Experimental Method**
9. Cite the lab manual where the original procedure was obtained. This would be a good "first sentence" for this section. Here is a sample sentence along with two ways of citing a reference:

a). The method as given by Crockford et al. Was followed, except that heating mantle was used in place of a micro-burner (1).

1. The method as given by Crockford et al. Was followed, except that heating mantle was used in place of a micro-burner.1
2. As shown above any modifications to the procedure in the reference should be noted.
3. Give a brief description of the procedure
4. This section is always written in the past tense (was, were, was used, etc.)
5. State the number of runs made and the conditions (concentrations, temperatures, etc.) at the end of this section.
6. **Results**
7. Tabulate the data that were measured. Tables should have a title with units, and explanatory captions. Do not exclude data if it seems out of line. The cause of this discordant data should be pointed out in the Discussion. Your final calculated results with error estimates should be listed in this section. Sometimes, data can be presented in form of graphs, in which case it is not necessary to tabulate it. A common example would be in a freezing point determination, a plot of time versus temperature would suffice for this data. Then report the actual freezing point from the graph in the results section.
8. Give a typical calculation (avoid the details). If the calculation is excessively long, it may be placed in the end in an Appendix.
9. Any graphs of data should be included in this section. Here are rules for graphs, especially if they are hand drawn.

A). Title the graph

B). Label the X and Y axis with units

C). Use "x's" to make the data points, not individual dots. Dots with circles, triangles, squares, etc. are fine.

D). Label any lines drawn in

E). The data points should cover at least half of the paper, especially for hand-drawn graphs.

1. **Discussion**

This is the most flexible section. It should always include:

1. A comparison of your results with literature values
2. Plausible reasons for discrepancies with the accepted values
3. Possible systematic errors and ways to reduce them
4. Improvements for the experimental technique
5. Special circumstances or difficulties that may have influenced the results
6. Suggestions of experiments to confirm suspicions of discrepancies of your results with accepted values
7. Suggestions of new/different chemical systems which may be interesting to try in the future
8. **References**

1. The **book** title can be either in italics or in quotations.

Garland, D. P.; Garland, C. W.; Nibler, J. W. *Experiments in Physical Chemistry*, 6th ed. McGraw-Hill: New York, **1996**.

1. A book with a plethora of information and an editor (like CRC Handbook of Chemistry and Physics):

*Handbook of Chemistry and Physics,* 72nd edition., R. C. Weast, editor, p F-10, CRC Press, Boca Raton, Florida, **1991**

1. A book with many authors (say 10 or more) with some editors:

Justice, J. C. In *Comprehensive Treatise of Electrochemistry,* Vol. 5, *Thermodynamic and Transport Properties of Aqueous and Molten Electrolytes*; Conway, B.E.; Bokris, J. O'M; Yeager, E. Eds.; Plenum: New York, **1983**; pp. 223-337.

1. A journal Article:

Gunning, H. E.; Gordon, A. R. *J. Chem. Phys*. **1942**, *10*, 126.

Here the year is in bold, the volume italicized, and the last number is the first page number. An option here is to underline the volume, and use regular type for both the year and the page number.

1. A journal Article with Title of the paper:

***"****Novel methods for the synthesis of functionalized indoles from arylhydroxylamines and activated acetylenes."* Hwu, J.R.; Patel, H.V.; Lin, R.J.; Gray, M.O. *J. Org. Chem.* **1994**, *59*, 1577-1582 and references cited therein.

**VIII. Fall 2019 CALENDAR**

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| Aug 16 | Residence Halls Open (Returning Students) |
| Aug 19 | Classes begin |
| Aug 19- Aug 23 | Late registration/Schedule Adjustment |
| Aug 30 | Campus Wide Assembly – First Year Students 9:00 am – Kean Hall |
| Sept 2 | Holiday-Labor Day |
| Sept 17 | Recognition of Constitution Day |
| Sept 22-28 | Student Study Week – No Activities Scheduled |
| Sept 30 – Oct 4 | Mid-term Examination Week-all classes meet as scheduled |
| Oct 11 | Last day to withdraw from a course and/or the University |
| Oct 15 | Founders Day (9:00 a.m. – Kean Hall) |
| Oct 19 | Homecoming |
| Oct 28 – Jan 17 | Registration for Spring 2020 |
| Nov 11 | Veterans’ Day |
| Nov 25-29 | Fall Break/Thanksgiving Holiday- No Classes |
| Nov 30 | Last Day of Class |
| Dec 2-6 | Final examinations |